

#### **ALEXANDER ROAD HIGH SCHOOL**

PHYSICAL SCIENCE CONTROL TEST MEMO

SEPTEMBER 2019

60 MIN

[10]

TOTAL = 60

JA, MH

GRADE 10

1.1 C ✓✓

**QUESTION 1: Multiple choice** 

- 1.2 A ✓✓
  1.3 D ✓✓
  1.4 C ✓✓
- 1.5
   D
   √√

   1.6
   B
   √√

   [12]

## **QUESTION 2:**

2.1	$N_2 + 3H_2 \rightarrow 2NH_3$ $\checkmark$ reactants $\checkmark$ products $\checkmark$ balancing	(3)
2.2	$(NH_4)_2SO_4 \checkmark \checkmark$	(2)
2.3	Mass cannot be created or destroyed in chemical reactions. $\checkmark$	(1)
2.4	$M_{reactants} = 2(14) + 6(1) = 34 \ g. \ mol^{-1} \checkmark$	(3)
	$M_{products} = 2(14 + 3(1)) = 34 \ g. \ mol^{-1} \checkmark$	
	∴ Mass is conserved. ✓	
2.5	Liquid phase 🗸	(1)

## **QUESTION 3:**

Volume, temperature, any other acceptable variable $\checkmark$	(1)
Direct proportionality ✓	(1)
The higher the concentration, the <u>more ions</u> $\checkmark$ there are in the solution and therefore there is a <u>greater movement of charge</u> $\checkmark$ resulting in a higher conductivity	(2)
LESS THAN $\checkmark$ , $Cu(OH)_2$ is insoluble $\checkmark$ and therefore will not form ions.	(2)
	Volume, temperature, any other acceptable variable $\checkmark$ Direct proportionality $\checkmark$ The higher the concentration, the <u>more ions</u> $\checkmark$ there are in the solution and therefore there is a <u>greater movement of charge</u> $\checkmark$ resulting in a higher conductivity. LESS THAN $\checkmark$ , $Cu(OH)_2$ is <u>insoluble</u> $\checkmark$ and therefore will not form ions.

# **QUESTION 4:**

4.1.1 Aqueous 
$$\checkmark$$
 (1)  
4.1.2 Solid  $\checkmark$  (1)  
4.2.1 The mass of one mole of a substance.  $\checkmark$  (1)  
4.2.2 Number of moles of solute per unit volume of solution.  $\checkmark$  (1)  
4.3.1  $M_{Na_2SO_4} = 2(23) + 32 + 4(16) = 142 g.mol^{-1} \checkmark \checkmark$  (answer only: 2/2) (2)  
4.3.2  $n = \frac{m}{M} \checkmark$   
 $n = \frac{28.4}{142}$  [ $\bigoplus$  -marking from 4.3.1]  
 $n = 0.2 mol \checkmark$   
 $c = \frac{n}{V} \checkmark$   
 $c = \frac{0.2}{0.25} \checkmark$   
 $\therefore c = 0.8 mol. dm^{-3} \checkmark$  (5)  
4.3.3  $n = \frac{m}{M}$   
 $0.2 \checkmark = \frac{m}{137 + 32 + 4(16)} \checkmark$   
 $0.2 = \frac{m}{233}$   
 $\therefore m = 46.6 g$  (3)  
4.3.4 % Yield  $= \frac{40}{46.6} \times 100 \checkmark$  [ $\bigoplus$  - marking from 4.3.3]

$$\therefore \%$$
 Yield = 85,84%  $\checkmark$  (answer only: 2/2) (2)

[16]

## **QUESTION 5:**

- 5.1 The change/difference in position.  $\checkmark \checkmark$  (2)
- 5.2 ✓ 5km vector (magnitude AND direction correct WITH label).
  ✓ 8km vector (magnitude AND direction correct WITH label).
- 5.3 ✓ Draws resultant vector from tail of first vector to head of last vector.



Magnitude: 8,7 km (allow 8,6 - 8,8 km)  $\checkmark$ (2)5.4 $65^{\circ}$  (allow 64 - 66) of resultant<br/>Direction to walk:  $245^{\circ}$  (180 + bearing of resultant)  $\checkmark$ (1)[7]

## **QUESTION 6:**

6.1 
$$T = \frac{1}{f} \checkmark = 0,025 \, s \checkmark$$

$$v = \frac{\Delta x}{\Delta t} \checkmark = \frac{0,038}{0.05} = 0,76 \, m.s \checkmark$$
(5)
6.2 
$$a = \frac{\Delta v}{\Delta t} \checkmark = \frac{0,28 - 0,76 \checkmark}{3 \times 0,025 \checkmark}$$

$$= -6,4 \, m.s^{-2} \text{ to the left or } 6,4 \, m.s^{-2} \text{ to the right }\checkmark$$
(4)

[9]

(2)